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assuming that no heat energy is lost from the calorimeter and given that the calorimeter has a negligible heat capacity the sum of the heat of dissolution  $q_{\text{soln}}$  and the change in heat energy of the urea water mixture must equal zero  $q_{\text{soln}} + mc\Delta T = 0$   
 $q_{\text{soln}} = mc\Delta T$   $m_{\text{soln}} = 5.13 \text{ g} + 91.95 \text{ g} = 97.08 \text{ g}$

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