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reaction rates get 3 of 4 questions to level up introduction to rate law learn rate law and reaction order worked example determining a rate law using initial rates data units of the rate constant practice introduction to rate law get 3 of 4 questions to level up concentration changes over time learn first order reactions

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4 answer the following questions regarding the kinetics of chemical reactions a the diagram below shows the energy pathway for the reaction  $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$  clearly label the following directly on the diagram i the activation energy  $E_a$  for the forward reaction ii  $\Delta H$  for the reaction b the reaction  $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$

## **8 3 kinetics and the mechanisms of reactions chemistry**

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questions to answer when a reaction releases energy where does the energy come from there is a rule of thumb that increasing the temperature by 10  $^{\circ}\text{C}$  will double the rate for many reactions what factor in the arrhenius equation is always changing explain why the reaction rate increases when the temperature increases

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three points one each for form of rate law  $\text{HgCl}_2$  exponent  $\text{C}_2\text{O}_4^{2-}$  2 exponent rate  $k$   $\text{HgCl}_2$   $\text{C}_2\text{O}_4^{2-}$  2 2 three points two for correct substitution and arithmetic one for units  $0.52 \times 10^4 \text{ M}^{-1} \text{ min}^{-1}$   $k = 0.0836 \text{ s}^{-1}$   $1.52 \times 10^2 \text{ M}^{-2} \text{ min}^{-1}$  note sometimes this unit is written as  $\text{l}^2 \text{ mol}^{-2} \text{ min}^{-1}$

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use the table below to answer the following questions find the order of the reaction ce a vs ce b find the overall order find the rate constant

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i do the reactants of the slow step match the species that are in the rate law ii does the stoichiometry of the slow step match the orders in the rate law iii do the individual steps of the mechanism add up to the overall stoichiometry of the equation

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kinetics gives information on the reaction rate and reaction mechanism the balanced chemical equation gives only the stoichiometry of the reaction

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answer the following questions related to the kinetics of chemical reactions  $\text{I}^- + \text{ClO}_2 + \text{OH}^- \rightarrow \text{IO}_3^- + \text{Cl}^-$  iodide ion is oxidized to hypoiodite ion  $\text{IO}_2^-$  by hypochlorite  $\text{ClO}_2^-$  in basic solution according to the equation above

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a the order of each reactant 1st order in a zero order in b b the order of the reaction overall reaction is 1st order c write the rate law rate  $k_a$  d calculate the rate constant  $k = 1.75 \times 10^{-2} \text{ s}^{-1}$  a the decomposition of  $\text{A}_2\text{B}_3$  is second order with a  $k = 6.5 \times 10^{-5} \text{ M}^{-1} \text{ s}^{-1}$  at 25 °C

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part ii select response select the best answer to each question 1 if the reaction  $2\text{HI} \rightarrow 2\text{H}_2 + \text{I}_2$  is first order which of the following will yield a linear plot a  $\log \text{HI}$  vs time b  $1/\text{HI}$  vs time c  $\text{HI}^2$  vs time d  $\ln \text{HI}$  vs time 2 consider the reaction  $\text{X} + \text{Y} \rightarrow \text{Z}$

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work in groups on these problems you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help

1 write the rate expression in terms of  $\Delta \text{reactant} / \Delta t$  and  $\Delta \text{product} / \Delta t$  for the reaction  $\text{N}_2\text{O}_3 \rightarrow \text{N}_2\text{O} + \text{O}_2$  number 2 determine

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