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idea identify three applications of hydrocarbons as a source of energy main and raw materials name an organic compound and explain what an organic chemist studies identify what each of the four molecular models highlights about a molecule compare and contrast saturated and unsaturated hydrocarbons

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22 1 hydrocarbons because carbon has four valence electrons carbon atoms always form four covalent bonds the carbon atoms in an alkane can be arranged in a straight chain or in a chain that has branches molecules of hydrocarbons such as alkanes are nonpolar molecules

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lesson summary organic chemistry and hydrocarbons hydrocarbons are carbon containing organic compounds that provide a source of energy carbon has four valence electrons so a carbon atom always forms four covalent bonds hydrocarbons are nonpolar molecules

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hydrocarbons contain only hydrogen and carbon atoms hydrocarbons in which each carbon atom is bonded to four other atoms are called alkanes or saturated hydrocarbons they have the general formula $C_n H_{2n+2}$ any given alkane differs from the next one in a series by a CH_2 unit

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hydrocarbons are organic compounds that contain only carbon and hydrogen the four general classes of hydrocarbons are alkanes alkenes alkynes and arenes aromatic compounds derive their names

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hydrocarbons are named using the iupac system although sometimes common names are used for certain familiar compounds all hydrocarbons are nonpolar molecules in general the fewer carbon atoms in a hydrocarbon the lower its melting and boiling points

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hydrocarbons are the important sources of energy the term hydrocarbon is self explanatory which means compounds of carbon and hydrogen only hydrocarbons play a key role in our daily life you must be familiar with the terms lpg and cng used as fuels

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hydrocarbons learning objectives by the end of this section you will be able to explain the importance of hydrocarbons and the reason for their diversity name saturated and unsaturated hydrocarbons and molecules derived from them describe the reactions characteristic of saturated and unsaturated hydrocarbons

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hydrocarbons identify hydrocarbons and the models used to represent them main idea hydrocarbons are carbon containing organic distinguish between saturated compounds that provide a source of energy and raw materials and unsaturated hydrocarbons

section 22 1 hydrocarbons pages 693 701

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this section describes how to identify and classify cyclic hydrocarbons it also explains the bonding in benzene cyclic hydrocarbons page 709 1 what is a cyclic hydrocarbon 2 the most abundant cyclic hydrocarbons contain or carbons 3 what are the names of the cyclic hydrocarbons represented below a

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hydrocarbons are organic compounds that contain only carbon and hydrogen the four general classes of hydrocarbons are alkanes alkenes alkynes and arenes aromatic compounds derive their names

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hydrocarbons are a group of organic chemicals made up only of the atoms carbon and hydrogen based on the amount of bonds between their carbon atoms hydrocarbons are divided into two categories saturated and unsaturated hydrocarbons hydrocarbons are used for several purposes the most essential as fuel

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hydrocarbons representative alkanes alkenes alkynes and aromatic compounds science 102 organic chemistry biochemistry module no 4 hydrocarbon chains and rings module overview many hydrocarbons occur in nature in addition to making up fossil fuels they are present in trees and plants as for example

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