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in the penny ante equilibrium a classroom activity chemtopic lab activity pennies are used as reactants and products in a reversible reaction to answer questions about the fundamental nature of equilibrium and what happens to the amount of reactants and products when it is reached

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the purpose of the lab was to see how equilibrium was changed with different reactants as well as product concentration when you increases the concentration of a reactant you re more likely getting a forward reaction while increasing the product gives reverse reactions post lab connection questions answer questions in complete sentences 1

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in this activity you will simulate a reaction that goes to equilibrium to use forward and reverse reactions to place a reaction in equilibrium to distinguish between the following terms reactants product forward reaction rate reverse reaction rate equilibrium and equilibrium constant

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this year i decided to tackle the conceptual understanding of equilibrium first with three different analogies and a process oriented guided inquiry learning pogil activity before even mentioning the words equilibrium expression

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in this lab you will explore the effect of le chatelier s principle on several chemical systems at equilibrium these are supplied in the theory section consider the third system you will study the aqueous ammonia solution

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what happens to the amount of reactants and products when equilibrium is reached what if more reactants or products are added to a system already at equilibrium in this activity pennies will be used as reactants and products in a reversible reaction to answer these questions and learn more about the fundamental nature of equilibrium

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because of its importance to analytical chemistry this chapter provides a review of equilibrium chemistry much of the material in this chapter should be familiar to you although some topics ladder diagrams and activity for example afford you with new ways to look at equilibrium chemistry

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for a given chemical system at equilibrium the concentrations of each reactant may differ but the value of  $k$  should be the same at a given temperature here s the best way to solve it expert verified 100 3 ratings a reaction said to be in equilibrium when rate of forward direction and

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