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an acid is a substance that yields an excess of hydrogen ions when dissolved in water there are three important points to understand about hydrogen in acids although all arrhenius acids contain hydrogen not all hydrogen atoms in a substance are capable of dissociating thus the ch 3 hydrogens of acetic acid are non acidic an

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an acid is any hydrogen containing substance that is capable of donating a proton hydrogen ion to another substance a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste

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two common indicators are litmus and phenolphthalein blue litmus turns red in the presence of an acid while phenolphthalein turns colorless acids

react with active metals to yield hydrogen gas recall that an activity series is a list of metals in descending order of reactivity

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learn the basics of acid base chemistry get the definitions and examples of acids and bases and learn about the neutralization reaction

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the earliest chemical definition of an acid the arrhenius definition says that an acid is a compound that increases the amount of hydrogen ion H^+ in aqueous solution an arrhenius base is a compound that increases the amount of hydroxide ion OH^- in aqueous solution

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what are acids acids are a molecule or other species which can donate a proton or accept an electron pair in reactions the word acid is derived from the latin word acidus which means sour

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acids have a ph value less than 7 pure water is neutral with a ph of 7 and bases have a ph greater than 7 here is a collection of useful and interesting facts about acids and bases the ph scale the ph scale is a way of measuring whether a liquid is an acid a base or neutral the scale runs from 0 strongly acidic to 14 strongly basic

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an acid is a substance that renders ionizable hydronium ion H_3O^+ in its aqueous solution it turns blue litmus paper red these dissociate in their aqueous solution to form their constituent ions as given by the following

examples based on their occurrence they are divided into two types

natural and mineral acids

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the pH concept a more concise way to express the acidity or basicity of

any aqueous solution important to know pH of any aqueous system

organic inorganic or biochemical

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acids acids are sour in taste turn blue litmus red and dissolve in water to

release H^+ ions example sulphuric acid H_2SO_4 acetic acid CH_3COOH nitric

acid HNO_3 etc properties of acids

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an acid is any hydrogen containing substance that is capable of donating

a proton hydrogen ion to another substance a base is a molecule or ion

able to accept a hydrogen ion from an acid acidic substances are usually

identified by their sour taste for more information on acids and bases

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the acid definition is given as any hydrogen that contains a substance capable of donating a proton a hydrogen ion to the other substance a

base is an ion or molecule that is able to accept a hydrogen ion from an acid usually the acidic substances are identified with their sour taste

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defining acids bases when acids are added to water they form positively charged hydrogen ions h the presence of h ions is what makes a solution acidic when alkalis are added to water they form negative hydroxide ions oh the presence of the oh ions is what makes the aqueous solution an alkali

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acids are a distinct class of compounds because of the properties of their aqueous solutions as outlined below aqueous solutions of acids are electrolytes meaning that they conduct electrical current some acids are strong electrolytes because they ionize completely in water yielding a great many ions

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to know the characteristic properties of acids and bases

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acids and bases are extremely important in many everyday applications our own bloodstream our environment cleaning materials and industry sulfuric acid is an economic indicator acid base theories arrhenius definition

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head of science conjugate acids bases a brønsted acid is a species that can donate a proton for example hydrogen chloride hcl is a brønsted acid as it can lose a proton to form a hydrogen h and chloride cl ion hcl aq h aq cl aq a brønsted base is a species that can accept a proton

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lecture notes for chapter 16 acids and bases i acids and bases a there are several ways to define acids and bases perhaps the easiest way to start is to list some of the properties of acids and bases b

notes on acids bases and p h acids bases and ph

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acids and bases are fundamental concepts in chemistry that describe different types of substances based on their properties and behavior in aqueous solutions the ph scale is used to measure the acidity or basicity of a solution

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